

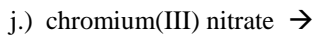
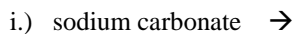
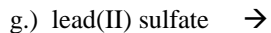
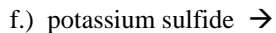
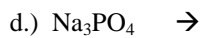
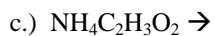
Name: _____

Solubility and Precipitation Reactions

1. On the basis of solubility rules indicate if the substance is (aq) soluble or (ppt) insoluble in water:

- | | |
|------------------------------------|--------------------------------------|
| a.) _____ Na_2SO_4 | k.) _____ $\text{Ni}(\text{OH})_2$ |
| b.) _____ ammonium phosphate | l.) _____ mercury(II) nitrate |
| c.) _____ Cu_3PO_4 | m.) _____ BaSO_4 |
| d.) _____ lead(II) sulfate | n.) _____ zinc acetate |
| e.) _____ AgOH | o.) _____ $\text{Cr}(\text{NO}_3)_6$ |
| f.) _____ mercury(I) chloride | p.) _____ lead(II) chloride |
| g.) _____ AgCl | q.) _____ Li_2CO_3 |
| h.) _____ magnesium sulfate | r.) _____ osmium(VIII) nitrate |
| i.) _____ Fe_2S_3 | s.) _____ $\text{Al}(\text{OH})_3$ |
| j.) _____ iron(II) carbonate | t.) _____ ammonium sulfide |

2. Write the products of dissociation for the following salts. If insoluble write *none*.



3. Complete the word equation then write a balanced chemical equation that includes phase notation. Circle the precipitate. If no reaction takes place write *No Reaction*

a.) Sodium chloride and silver nitrate →

b.) Phosphoric acid and copper(II) sulfate →

c.) sodium chloride and magnesium bromide →

d.) sodium sulfide and iron(III) nitrate →

e.) ammonium phosphate and tin(II) nitrate →

f.) lead(II) nitrate and sulfuric acid →

g.) silver acetate and sodium sulfide →

h.) ammonium carbonate and sodium phosphate →

i.) ammonium phosphate and lead(II) acetate →