

## In Class Pre Test Practice: Limiting Reactant Stoichiometry

1. For the reaction of 50.00 grams of  $\text{NH}_3$  with 150.00 grams of  $\text{HCl}$ ;  $\text{NH}_3 + \text{HCl} \longrightarrow \text{NH}_4\text{Cl}$   
a.) Which reactant is limiting. Show your work.



- b.) What mass of  $\text{NH}_4\text{Cl}$  should be produced?

2. 2.50 grams of sodium phosphate reacts with 2.50 grams of cobalt(II) chloride.  
a.) Write a balanced equation for this reaction:

- b.) Which reactant is limiting. Show your work.

- c.) What mass of cobalt(II) phosphate would be produced?

- d.) What mass of excess reactant is wasted?