Chemistry PracTest: Classifying, Predicting. And Writing Balanced Equations

a. _____ What gas from our *Stations Lab* extinguishes a flame?

b. _____ How many atoms of H are indicated by: $3(NH_4)_2SO_4$

c. _____ What ions do acids produce?

d. _____ What ions do bases produce?

e. _____ How many electrons does each Br atom gain when it reacts with a metal?

f. _____ Is the Br atom oxidized or reduced?

g. _____ How many electrons does each Al atom lose when it reacts with a non-metal?

h. _____ Is the Al atom oxidized or reduced?

i. _____ and _____ Products formed when an acid reacts with a base.

j. _____ Ammonia is a(n) (Acid / Base).

k. _____ _____ Evidence of a double displacement reaction is the formation of a _____ or _____.

l. _____ AgCl is placed in water and sinks to the bottom. What phase notation should the AgCl have?

m. _____ Reaction type in which a compound reacts with elemental oxygen and releases a large amount of heat.

n. _____ 10.0 g of ZnI_2 is decomposed. If 2.0 grams of Zn is produced how many grams of I_2 are also produced?

Write a balanced chemical equation from the given information. Focus on the reactants which are <u>underlined</u>.

a. A piece of <u>zinc</u> metal is added to a solution of <u>hydrochloric acid</u>.

b. An electric current is passed through a large jar of <u>water</u> and bubbles are produced.

c. Rubbing alcohol, $\underline{C_3H_8O}$, is combusted.

d. A pea size piece of <u>rubidium</u> metal is added to a beaker of <u>water</u> , H-OH).

e. A solution of <u>rubidium phosphate</u> is added to a solution of <u>nickel(II) nitrate</u>.

f. A solution of <u>silver nitrate</u> is added to a piece of coiled <u>copper</u> metal in a test tube.