

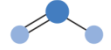



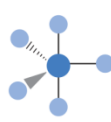







Electron Groups	Bonding Groups	Lone Pairs	Electron Geometry	Molecular Geometry	Approximate Bond Angles	Example
2	2	0	Linear	Linear	180°	$\text{:}\ddot{\text{S}}=\text{C}=\ddot{\text{S}}\text{:}$ 
3	3	0	Trigonal planar	Trigonal planar	120°	$\begin{array}{c} \text{:Cl:} \\ \\ \text{:Cl}-\text{B}-\text{Cl:} \\ \\ \text{:Cl:} \end{array}$ 
3	2	1	Trigonal planar	Bent	<120°	$\text{:}\ddot{\text{O}}=\ddot{\text{S}}-\ddot{\text{O}}\text{:}$ 
4	4	0	Tetrahedral	Tetrahedral	109.5°	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$ 
4	3	1	Tetrahedral	Trigonal pyramidal	<109.5°	$\begin{array}{c} \text{H}-\ddot{\text{P}}-\text{H} \\ \\ \text{H} \end{array}$ 
4	2	2	Tetrahedral	Bent	<<109.5°	$\text{H}-\ddot{\text{S}}-\text{H}$ 
5	5	0	Trigonal bipyramidal	Trigonal bipyramidal	120° (equatorial) 90° (axial)	$\begin{array}{c} \text{:Cl:} \\ \\ \text{Cl}-\text{P}-\text{Cl:} \\ \\ \text{:Cl:} \end{array}$ 
5	4	1	Trigonal bipyramidal	Seesaw	<120° (equatorial) <90° (axial)	$\begin{array}{c} \text{:F:} \\ \\ \text{F}-\text{S}-\text{F:} \\ \\ \text{:F:} \end{array}$ 
5	3	2	Trigonal bipyramidal	T-shaped	<90°	$\begin{array}{c} \text{:F:} \\ \\ \text{F}-\text{Cl}-\text{F:} \\ \\ \text{:F:} \end{array}$ 
5	2	3	Trigonal bipyramidal	Linear	180°	$\text{:F}-\text{Xe}-\text{F:}$ 
6	6	0	Octahedral	Octahedral	90°	$\begin{array}{c} \text{:F:} \\ \\ \text{F}-\text{S}-\text{F:} \\ \\ \text{:F:} \end{array}$ 
6	5	1	Octahedral	Square pyramidal	<90°	$\begin{array}{c} \text{:F:} \\ \\ \text{F}-\text{Br}-\text{F:} \\ \\ \text{:F:} \end{array}$ 
6	4	2	Octahedral	Square planar	90°	$\begin{array}{c} \text{:F:} \\ \\ \text{F}-\text{Xe}-\text{F:} \\ \\ \text{:F:} \end{array}$ 